



- How do we know how many of each element we need to write the correct formula? (think Lewis diagrams or PPCM)

Use the elements on the previous page and the ions they made to show how they combine. There are 9 possible compounds!

- Finally, name the ionic compound in French (A great opportunity for you to laugh at your parents' French accent.)

Rules for naming

- Try some examples!

Li and F

Ca and Br

Al and P

Mg and Cl

Ga and O

- A little assignment of four questions to assess their learning. **FILM THIS AND THEN SEND IT TO ME BY ATTACHING IT TO THE CHAT... Do I have their permission to post the video to the website? YES / NO**

## Evaluation of ionic compounds

1. On the periodic table, identify where we find the metals and the non-metals. /2

PERIODIC TABLE OF ELEMENTS

1 H Hydrogen Nonmetal	2 He Helium Noble Gas																
3 Li Lithium Alkali Metal	4 Be Beryllium Alkaline Earth Metal	5 B Boron Metalloid	6 C Carbon Nonmetal	7 N Nitrogen Nonmetal	8 O Oxygen Nonmetal	9 F Fluorine Halogen	10 Ne Neon Noble Gas										
11 Na Sodium Alkali Metal	12 Mg Magnesium Alkaline Earth Metal	13 Al Aluminum Metal	14 Si Silicon Metalloid	15 P Phosphorus Nonmetal	16 S Sulfur Nonmetal	17 Cl Chlorine Halogen	18 Ar Argon Noble Gas										
19 K Potassium Alkali Metal	20 Ca Calcium Alkaline Earth Metal	21 Sc Scandium Transition Metal	22 Ti Titanium Transition Metal	23 V Vanadium Transition Metal	24 Cr Chromium Transition Metal	25 Mn Manganese Transition Metal	26 Fe Iron Transition Metal	27 Co Cobalt Transition Metal	28 Ni Nickel Transition Metal	29 Cu Copper Transition Metal	30 Zn Zinc Transition Metal	31 Ga Gallium Metal	32 Ge Germanium Metalloid	33 As Arsenic Metalloid	34 Se Selenium Nonmetal	35 Br Bromine Halogen	36 Kr Krypton Noble Gas
37 Rb Rubidium Alkali Metal	38 Sr Strontium Alkaline Earth Metal	39 Y Yttrium Transition Metal	40 Zr Zirconium Transition Metal	41 Nb Niobium Transition Metal	42 Mo Molybdenum Transition Metal	43 Tc Technetium Transition Metal	44 Ru Ruthenium Transition Metal	45 Rh Rhodium Transition Metal	46 Pd Palladium Transition Metal	47 Ag Silver Transition Metal	48 Cd Cadmium Transition Metal	49 In Indium Metal	50 Sn Tin Metal	51 Sb Antimony Metalloid	52 Te Tellurium Metalloid	53 I Iodine Halogen	54 Xe Xenon Noble Gas
55 Cs Cesium Alkali Metal	56 Ba Barium Alkaline Earth Metal	57 La Lanthanum Lanthanide	58 Ce Cerium Lanthanide	59 Pr Praseodymium Lanthanide	60 Nd Neodymium Lanthanide	61 Pm Promethium Lanthanide	62 Sm Samarium Lanthanide	63 Eu Europium Lanthanide	64 Gd Gadolinium Lanthanide	65 Tb Terbium Lanthanide	66 Dy Dysprosium Lanthanide	67 Ho Holmium Lanthanide	68 Er Erbium Lanthanide	69 Tm Thulium Lanthanide	70 Yb Ytterbium Lanthanide	71 Lu Lutetium Lanthanide	
87 Fr Francium Alkali Metal	88 Ra Radium Alkaline Earth Metal	89 Ac Actinium Actinide	90 Th Thorium Actinide	91 Pa Protactinium Actinide	92 U Uranium Actinide	93 Np Neptunium Actinide	94 Pu Plutonium Actinide	95 Am Americium Actinide	96 Cm Curium Actinide	97 Bk Berkelium Actinide	98 Cf Californium Actinide	99 Es Einsteinium Actinide	100 Fm Fermium Actinide	101 Md Mendelevium Actinide	102 No Nobelium Actinide	103 Lr Lawrencium Actinide	

2. Draw the Lewis dot structure for the following elements; /6

Be                  Cl                  As                  Al                  K                  S

3. Combine the following elements to create the correct ionic compound and name them. /8

K and S                  Be and Cl                  Al and Cl                  Al and S

How did we do? 13-16 pts – FULL NERD STATUS  
10-12 pts – PRETTY DANG NERDY  
6-10 pts – NERD IN TRAINING  
1-6 pts – DEFINITELY THE TEACHER'S FAULT!

***Parents,***

***Thank you for participating and I hope you had some fun seeing what we are learning in this unit.***

***Just wait until you get to balance chemical equations!***